ketone **8** was identified. Furthermore small quantities of the furan **9** were observed in the acylation of *n*-butylacetylene, but **9** could be shown to arise from isomerization of the allene **7**.

The somewhat surprising failure to isolate any of the substitution products **6** upon reaction with **4** suggests a significant change in some energy parameter in going from the acyclic acyl to cycloacyl cations. While the enhanced strain in forming a fused ring system may be important, the fact that the intermediate carbenium ion C in the present case is secondary, whereas in the substitution reactions with the acyl tetrafluoroborates it would have to be primary, is also undoubtedly significant. Studies are presently underway to examine these various possibilities.

References and Notes

(1) Portions of this work were reported at 169th National Meeting of the American Chemical Society, Philadelphia, Pa., April 6–11, 1975, and 9th Great Lakes Regional Meeting of the American Chemical Society, St. Paul, Minn., June 4–6, 1975.

(2) A. A. Schegolev, W. A. Smit, G. V. Roitburd, and V. F. Kucherov, *Tetrahedron Lett.*, 3373 (1974).

(3) The acyl cation is undoubtedly linear: G. A. Olah and M. B. Comisarow, *J. Am. Chem. Soc.*, **88**, 4442 (1966).

(4) G. Olah et al., *J. Am. Chem. Soc.*, **95**, 6827 (1973), and references cited therein.

(5) D. M. Brower and H. Hogeveen, *Prog. Phys. Org. Chem.*, 9, 179 (1972).

(6) For recent reviews on vinyl cation chemistry see (a) M. Hanack, *Acc. Chem. Res.*, **3**, 209 (1970); (b) G. Modena and U. Tonellato, *Adv. Phys. Org. Chem.*, **9**, 185 (1971); (c) P. J. Stang, *Prog. Phys. Org. Chem.*, **10**, 205 (1973); (d) L. R. Subramanian and M. Hanack, *J. Chem. Educ.*, **52**, 80 (1975).

(7) Hydride shifts of this order have been known for a long time in transannular reactions of trivalent carbenium ions; for a review see E. L. Eliel, "Stereochemistry of Carbon Compounds", McGraw-Hill, New York, N. Y., 1962, Chapter 9.

(8) Origin of chlorine is abstraction from solvent: see W. S. Johnson, M. B. Gravestock, R. J. Parry, and D. A. O'Korie, *J. Am. Chem. Soc.*, **94**, 8604 (1972).

(9) This product corresponds to the addition of Friedel-Crafts acid chloride-$AlCl_3$ complexes to acetylenes; see H. Martens, F. Janssens, and G. Hoornaert, *Tetrahedron*, **31**, 177 (1975), and references cited therein.

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Additions and Corrections

The Laser Initiated Oxidative Photoaddition of *p*-Benzoquinone to Cyclooctatetraene [*J. Am. Chem. Soc.*, **96**, 2955 (1974)]. By R. Marshall Wilson,* Edward J. Gardner, R. C. Elder, Richard H. Squire, and L. Roxane Florian, Department of Chemistry, University of Cincinnati, Cincinnati, Ohio 45221.

Owing to an inadvertent interchange of data files, the final cycles of least-squares refinement of the crystal structure of **13** were performed on data not corrected for absorption. We anticipate that little, if any bias has been introduced into the reported atom positions or the bond lengths and angles derived therefrom. On the other hand, the anisotropic temperature parameters reported in Table II and the root-mean-square displacements in Table V are undoubtedly systematically biased by this error and should not be taken to describe the vibrational amplitudes in the crystalline sample of **13**.

Stereochemistry of Cobalt Porphyrins. I. The Structure and Characterization of 2,3,7,8,12,13,17,18-Octaethylporphinatobis(3-methylpyridine)cobalt(II) [*J. Am. Chem. Soc.*, **96**, 4440 (1974)]. By Robert G. Little and James A. Ibers,* Department of Chemistry, Northwestern University, Evanston, Illinois 60201.

The value of γ should be 66.72° rather than 113.28°.

Stereochemistry of Cobalt Porphyrins. II. The Characterization and Structure of *meso*-Tetraphenylporphinatobis(imidazole)cobalt(III) Acetate Monohydrate Monochloroformate, $[(Co(Im)_2(TPP)][OAc]\cdot H_2O\cdot CHCl_3$ [*J. Am. Chem. Soc.*, **96**, 4447 (1974)]. By Joseph W. Lauher and James A. Ibers,* Department of Chemistry and Materials Research Center, Northwestern University, Evanston, Illinois 60201.

The value of α should be 103.94° rather than 105.94°.

The following corrections should be made in Table III:
z for H of PB-C(4) should be −0.016 rather than 0.157.
z for PD-C(6) should be 0.5446 (12) rather than p.5446 (12).
y for H-C(2) should be −0.231 rather than 0.231.
y for H-C(3) should be −0.111 rather than 0.111.

Matrix Photolysis of 1,2,3-Thiadiazole. On the Possible Involvement of Thiirene [*J. Am. Chem. Soc.*, **96**, 6768 (1974)]. By A. Krantz* and J. Laureni, Department of Chemistry, State University of New York, Stony Brook, New York 11790.

The sentence beginning on the sixth line in the left-hand column on page 6769 should read, "The assignment of these bands to a single species is most clearly evident upon irradiation with light of $\lambda > 220$ nm, which results in loss of the thioketene spectrum and the appearance of a band at 1520 cm^{-1}, perhaps due to a secondary product from carbon monosulfide."

A Sensitive Probe for Double Layer Structure. Potential Dependent Competitive Cyanation and Methoxylation of 1,4-Dimethoxybenzene [*J. Am. Chem. Soc.*, **97**, 1499 (1975)]. By N. L. Weinberg,* D. H. Marr, and C. N. Wu, Hooker Chemical and Plastics Corporation, Research Center, Grand Island, New York 14302.

Page 1501 (under eq 4): A solution of 0.5 M $NaCN/CH_3OH$ will actually contain approximately 0.48 M cyanide and 0.02 M methoxide ions.

Page 1503 (left-hand column, line 4): Case II, $k_3/k_4 = 1/40$, and the ratio (CN^-/CH_3O^-) at $E = 1.35$ V is 0.48/0.02. Here an arbitrary potential is chosen midway in the potential region under consideration at which point the relative bulk concentration of nucleophiles is assumed to be equal to that available for reaction at the electrode. Significantly, a potential less than about $E = 1.15$ V cannot be se-